
The Activity Series Worksheet Answers

an activity series - chemdemos - an activity series continued 2 21 linn scientific inc all rights reserved 1c1: many properties of atoms exhibit periodic trends that are reflective of the periodicity of electronic structure. chemical reactions can be classified by considering what the reactants are, what the products are, or how they change from one **activity series lab - auburnnet** - activity series lab (microscale) purpose: to study the chemical activity of common metals safety precautions wear goggles at all times. review precautions for handling acids. since silver nitrate may stain skin and clothing, avoid contact. procedure (part i): 1. sketch out a matrix to record your data. along the top of the matrix, you should list **teaching transparency worksheet the activity series** - teaching transparency worksheet the activity series use with chapter 9, section 9.2 1. for each of the following pairs of elements, underline the one that would replace the other element in a compound. a. calcium, tin e. iron, copper b. bromine, fluorine f. iodine, chlorine c. aluminum, potassium g. silver, lead d. zinc, sodium 2. **developing the activity series - asminternational** - developing the activity series. overview: the students will develop their own activity series of metals based on lab results. qualitative observations will be used. tie-ins include single-replacement reactions (srr), oxidation-reduction, practical applications such as galvanization, and predicting reactions. students work in **the activity series for single displacement reactions** - the activity series for single displacement reactions chem 30a h a more active element (higher on the table above) will form an ion and displace the ion of like charge from the compound. the displaced ion will become a neutral element (see diatomic elements below). **chemistry single replacement reaction worksheet** - using the activity series table, complete the following reactions by writing the products that are formed. be sure to balance each equation. if no single replacement reaction occurs, write nr to the right of the arrow. 1. $ag + kno_3$ 2. $zn + agno_3$ 3. $al + h_2so_4$ 4. $cl_2 + ki$ 5. $li + h_2o$ 6. $cu + feso_4$ 7. $na + h_2o$ 8. **the activity series - weebly** - the activity series 3 7. an active metal readily gives up its electron(s) and is oxidized. there are four metals involved in the reactions of model 1—magnesium, aluminum, copper, and zinc. rank these metals from **name using the activity series date: - mychemistry** - 7sed on the activity series,, which ion will oxidize pb to pb²⁺? 1)magnesium 2)nickel 3)copper 4)silver 8. which metal reacts spontaneously with a solution containing zinc ions? **experiment 8 activity series (single displacement reactions)** - experiment 8: activity series (single displacement reactions) purpose a) reactions of metals with acids and salt solutions b) determine the activity of metals c) write a balanced molecular equation, complete ionic equation, and net ionic equation for single **solubility rules and activity series** - solubility rules and activity series compounds with these ions are soluble (aq) exceptions that are solid (s) group i metal ions (na⁺, k⁺, li⁺, etc.) none (always aqueous) +ammonium ion (nh₄⁺) none (always aqueous) - nitrate ion (no₃⁻) none (always aqueous) -acetate ion (c₂h₃o₂⁻) none (always aqueous) **activity series of metals - mr. f's classroom & laboratory** - the activity series of the metals is an invaluable aid to predicting the products of replacement reactions. it also can be used as an aid in predicting products of some other reactions. pay attention to the notes below as they are provided to help you make better use of the activity series than just the list of metals by themselves. 1.

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